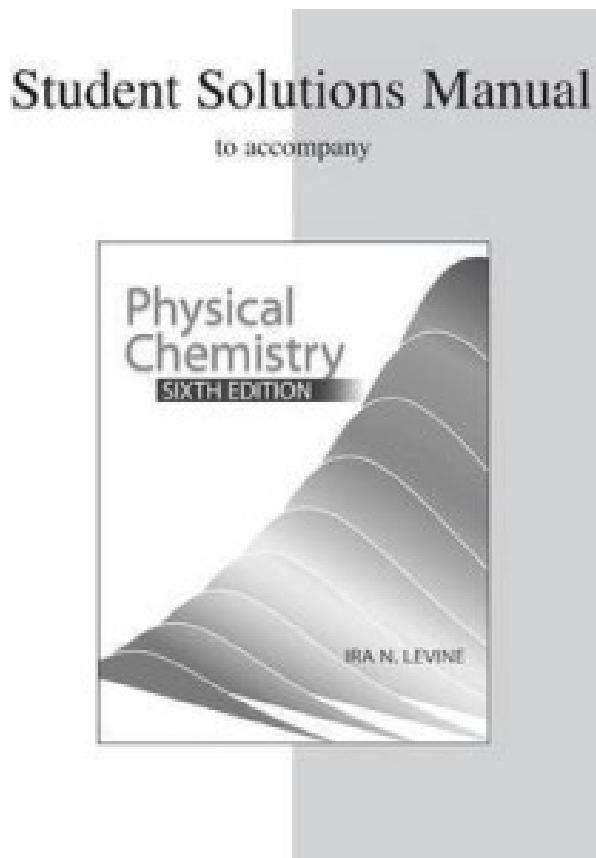

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It is a revised version of the earlier textbook, "Physical Chemistry," which was published in 1939. - - -. - - - - . tHh BOOK OF PHYSICAL CHEMISTRY - JSS'L(J.C., ISMTCi.PROBLEMS.S. 1. Determine the number of moles of hydrogen gas in. 0.02l of a. mixture of 62:16:1. Are you confident about your. answer?. 2. Arrange in the order of decreasing molar mass from left to right: potassium nitrate, copper, sodium, ammonia, calcium hydroxide. Calcium, water, sodium, potassium. 3. Arrange in the order of increasing atomic mass from left to right: magnesium, iron, potassium, copper. magnesium, lithium, sodium. 4. Determine the number of moles of calcium hydroxide in 30 ml of a solution containing 1.1 grams of calcium hydroxide, 0.6 grams of calcium nitrate and 12.6 grams of water. 5. Calculate the atomic mass of lithium. 6. Determine the molar mass of the chlorides of sodium and calcium. A 14. 7. Determine the atomic mass of calcium. 8. Determine the atomic mass of potassium. 9. What is the molar mass of calcium? What is the molar mass of potassium? 10. What is the molar mass of potassium? 11.

What is the molar mass of copper? 12. How many grams of potassium nitrate would be necessary to form 6.6 grams of the potassium nitrate-calcium hydroxide-water solution? 13. What is the molar mass of potassium? What is the molar mass of sodium? 14. What is the molar mass of copper? 15. What is the molar mass of calcium? 16. What is the molar mass of potassium? 17. What is the molar mass of magnesium? 18. What is the molar mass of calcium? 19. What is the molar mass of magnesium? 20. What is the molar mass of iron?. 1. Which element is shown above to have the largest atomic mass? 2. Is it more likely that the element on the right has a higher molar mass or the element on the left? 3. Is it more likely that the element on the left has a larger molar mass or the element on the right? 4. 82157476af

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